

"NOVEL ORGANIC/INORGANIC-OXIDE MULTILAYER MATERIALS"**FIELD OF INVENTION**

5 The present invention relates to novel organic/inorganic-oxide multilayer materials. More particularly but not exclusively the invention relates to novel organic/inorganic-oxide materials based on single or multiple atomic layers of tungsten oxide, molybdenum oxide or other metal oxides interspersed between organic spacer layers.

10 BACKGROUND

The development of organic conductors and semiconductors in recent decades has opened the door to the possibility of flexible plastic displays and low-cost logic and storage devices [see for example P.Yam, 'Plastics get wired.' *Scientific American*, 273 (1), 74-79 (July 1995), or J.C. Scott, 'Conducting polymers: From novel science to new technology.' *Science*, 278 (5346), 2071-2072 (12 Dec 1997).]. There is much international effort to push these materials through to practical devices [see for example the above noted references or: J. S. Miller, 'Conducting Polymers – materials of commerce', *Advanced Materials*, 5 (7/8 & 9), 587-589 & 671-676, (1993).]. The overwhelming significance of these developments has been recognised by the award of the Nobel Prize in Chemistry for the year 2000 to MacDiarmid, Heeger and Shirakawa, the discoverers of the conducting polymer polyacetylene [see for example R. F. Service 'Getting a Charge Out of Plastics' *Science*, 290, 425-427, (20 October 2000)]. However, a significant limitation in the use of organic semiconductors is their low electron mobility arising from the weak van der Waals bonding between neighbouring polymer molecules. This, in turn, limits device switching speeds. High electron mobility can be found in doped crystalline inorganic oxides (because of their strong covalent bonding of their extended framework) but these materials demand high temperatures for synthesis and are brittle, non-flexible materials.

A specific example of such an oxide material is tungsten trioxide, WO₃. This material adopts a wide range of electronic properties ranging from an n-type semiconductor to a metal depending upon doping level. These doping states are accompanied by the well-known changes in colour observed in the so-called tungsten bronzes that have been used, for example, in pigments for paint. More advanced

applications include architectural glazings, ophthalmic devices, instrumentation devices and displays. Tungsten trioxide has a perovskite ABO_3 structure with the usually 12-coordinated A-sites vacant, consistent with the 6+ valence state of W^{6+} . Each W ion is 6-fold coordinated to oxygen atoms in a roughly octahedral coordination. The octahedra are 5 corner shared as shown by the schematic section in Figure 1 where the octahedra are represented by the corner shared diamonds. Electron doping into the conduction band is achieved by inserting additional cations into the otherwise vacant 12-coordinated A sites. Molybdenum trioxide, MoO_3 , exhibits much the same chemistry and structure. (While ABO_3 is the usual generalisation of the perovskite formula the identifier "A" will be used 10 in the following to describe general organic groups. To avoid confusion the general formula ZMO_3 will be used in place of ABO_3 where Z refers to the 12-coordinated perovskite site and M refers to the 6-coordinated approximately octahedral site).

OBJECT OF THE INVENTION

15 It is an object of the present invention to provide novel materials and methods for preparing such materials, exhibiting electron mobility/conductivity characteristics. It is an additional/alternative object to provide a novel material for application as a semiconductor and/or superconductor. It is an additional/alternative object to provide the public with a useful alternative.

SUMMARY OF INVENTION

In a first aspect of the invention there is provided a layered organic-inorganic oxide material comprising or including:

- 25
- one or more layers of metal oxide consisting of one or more atomic planes of corner-shared MO_6 octahedra, where M is the metal, and
 - one or more layers of organic molecules,
wherein the metal-oxide layers alternate with one or more organic layers alternate to form a periodic planar structure.

Preferably, the metal M is W, V or Mo, or a combination of these.

30 Preferably the metal M is metal M is W, V or Mo, or a combination of these and wherein other high valency metals such as Ti, Nb, Ta, Ru and Re are used in partial combination with Mo, W or V.

Preferably the material a general formula $X.M_mO_{3m+1}$ wherein M is the metal, and X is an organic cation and m=1, 2, 3.

Preferably the organic cation is bidentate.

Preferably the configuration of organic layer relative to the inorganic layer is 5 eclipsed.

In one form the organic cation is a diammonium cation, the material is of composition $NH_3.A.NH_3.M_mO_{3m+1}$.

Preferably m=1, such that each inorganic oxide atomic plane alternates with an 10 organic layer; alternatively m=2, the composition is $NH_3.A.NH_3.M_2O_7$ and wherein the organic oxide exists as a double atomic plane layer of corner shared MO_6 octahedra, such that the material has the stacking structure -A-O-MO₂-O-MO₂-O-A.

Preferably the organic cation is an aliphatic diammonium cation, and A = $(CH)_n$, n= 1,2,.....

Preferably on the organic cation the ammonium cation groups are positioned on the 15 terminal alkane units of A.

In another form the material has the chemical formula $NH_3(CH_2)_nNH_3MO_4$.

Preferably with n=2 or 6 or 12.

Preferably the organic cation is an aromatic diammonium cation.

Preferably A = C₆H₄ and the organic cation is NH₃C₆H₄NH₃.

Alternatively the organic cation comprises an aromatic ring with two aliphatic 20 side chains of equal or unequal length each side chain terminated by an ammonium ion, the organic cation having the general formula $NH_3(CH_2)_nC_6H_4(CH_2)_mNH_3$.

Preferably adjacent aromatic rings are crosslinked to form an organic polymer layer.

Preferably the organic polymer layer is conducting.

In another form the material has a general formula $X'^2.M_mO_{3m+1}$ wherein M is the metal, and X' is an organic cation and m=1, 2, 3.

Preferably the organic cation is monodentate.

Preferably the configuration of organic layer relative to the inorganic layer is 30 staggered.

Preferably both organic cations are monoammonium cations and the material is of composition $(NH_3.A')_2.M_mO_{3m+1}$.

Preferably m=1, such that each inorganic oxide atomic layer alternates with an organic layer.

Alternatively m=2, the composition is $(\text{NH}_3.\text{A}')_2.\text{M}_2\text{O}_7$ and wherein the organic oxide exists as a double atomic plane layer having approximately the ZWO_3 perovskite structure with the Z sites vacant such that the material has the stacking structure $\text{NH}_3.\text{A}'\text{-MO}_2\text{-O- MO}_2\text{-A'}\text{-NH}_3$.

Preferably one or both organic cation is an aliphatic ammonium cation, and $\text{A}' = (\text{CH})_n$, $n = 1, 2, \dots$

Alternatively one or both organic cation is an aromatic ammonium cation.

Preferably the aromatic ring has a side chain which is aliphatic and terminated by an ammonium ion, having the formula $(\text{C}_6\text{H}_5(\text{CH}_2)_m\text{NH}_3)_2\text{MO}_4$ where $m=0, 1, 2, 3, \dots$

Preferably adjacent aromatic rings are crosslinked to form an organic polymer layer.

Preferably the organic polymer layer is conducting.

Preferably dopants are introduced into the structure.

Preferably the dopant is selected from one or more of an alkali cation a methylammonium cation, replacing ammonium groups, field-effect injected electrons or field-effect injection electron holes.

Preferably the dopant is present in the inorganic oxide layers and the doping state of the oxide is adjusted such that the oxide exhibits superconductivity above the temperature of 40K.

Preferably the doping state of the oxide is adjusted such that the oxide exhibits superconductivity above the temperature of 90K.

Preferably M is partially or fully substituted by a magnetic transition metal ion so as to display magnetically ordered states.

In a further embodiment there is provided an organic/inorganic oxide material of any of claims 1 to 33 in which the oxide layer comprising MO_4 , M_2O_7 or $\text{M}_m\text{O}_{3m+1}$ is wholly replaced by any of the following oxide layers CuO_2 , NiO_2 , CoO_2 , $\text{CuO}_2\text{CaCuO}_2$, $\text{Ca}_{m-1}\text{Cu}_m\text{O}_{2m}$, $m=1, 2, 3, \dots$, $\text{NiO}_2\text{CaNiO}_2$, $\text{Ca}_{m-1}\text{Ni}_m\text{O}_{2m}$, $m=1, 2, 3, \dots$, square pyramidal MnO_3 , square pyramidal RuO_3 , octahedral RuO_4 , $\text{O-MnO}_2\text{-Y-MnO}_2\text{-O}$, $\text{O-MnO}_2\text{-Ca-MnO}_2\text{-O}$, $\text{O-RuO}_2\text{-Y-RuO}_2\text{-O}$, or $\text{O-RuO}_2\text{-Ca-RuO}_2\text{-O}$.

According to a further aspect of the invention there is provided a layered organic-inorganic oxide material comprising or including:

- One or more layers of metal oxide consisting of one or more atomic planes of metal oxide having substantially the ZMO_3 perovskite structure ($M=$ metal) with the Z sites vacant, and
- one or more layers of organic molecules,
wherein the metals form divalent cations and are coordinated into a corner-shared square-planar structure, or wherein the metals form tetravalent cations and are coordinated into a corner-shared square-pyramid structure,
wherein one or more metal-oxide layers alternate with one or more organic layers alternate to form a periodic planar structure.

Preferably the metal, M , is Cu, Ni, Ru, Mn, or a combination of these.

Preferably higher order structures are formed with two or more oxide layers each separated by an alkali earth ion which is situated in the perovskite A-site.

Preferably the alkali earth ion is calcium.

Preferably the material has the general formula of one of: $NH_3.A.NH_3CuO_2$, $(A.NH_3)_2CuO_2$, $NH_3.A.NH_3Ca_{m-1}Cu_mO_{2m}$, $m=1, 2, 3, \dots$, $(A.NH_3)_2Ca_{m-1}Cu_mO_{2m}$, $m=1, 2, 3, \dots$, $NH_3.A.NH_3NiO_2$, $(A.NH_3)_2NiO_2$, $NH_3.A.NH_3Ca_{m-1}Ni_mO_{2m}$, $m=1, 2, 3, \dots$, $(A.NH_3)_2Ca_{m-1}Ni_mO_{2m}$, $m=1, 2, 3, \dots$, and $NH_3.A.NH_3MnO_3$, $(A.NH_3)_2MnO_3$, $NH_3.A.NH_3Ca_{m-1}Mn_mO_{2m+2}$, $m=1, 2, 3, \dots$, $(A.NH_3)_2Ca_{m-1}Mn_mO_{2m+2}$, $m=1, 2, 3, \dots$, $NH_3.A.NH_3RuO_3$, $(A.NH_3)_2RuO_3$, $NH_3.A.NH_3Ca_{m-1}Ru_mO_{2m+2}$, $m=1, 2, 3, \dots$, $(A.NH_3)_2Ca_{m-1}Ru_mO_{2m+2}$, $m=1, 2, 3, \dots$

Preferably dopants are introduced into the structure.

Preferably the dopant is selected from one or more of an alkali cation a methylammonium cation, replacing ammonium groups, field-effect injected electrons or field-effect injection electron holes.

Preferably the dopant is present in the inorganic oxide layers and the doping state of the oxide is adjusted such that the oxide exhibits superconductivity above the temperature of 40K.

Preferably the doping state of the oxide is adjusted such that the oxide exhibits superconductivity above the temperature of 90K.

Preferably M is partially or fully substituted by a magnetic transition metal ion so as to display magnetically ordered states.

According to a further aspect of the invention there is provided a method of preparing a layered inorganic-organic material which comprises or includes:

- 5 - one or more layers of metal oxide, and
 - one or more organic layers,

wherein the layers exist substantially in a perovskite structure, the method comprising or including the steps of contacting a source of metal and/or oxide with a source of the organic layer such that the layer structure substantially self 10 assembles.

Preferably the material is of the general structure $\text{NH}_3\cdot\text{A}\cdot\text{NH}_3\cdot\text{M}_m\text{O}_{3m+1}$ and is prepared either:

- by reaction of a diaminoalkane with tungstic acid (when the metal is W) or molybdic acid (when the metal is Mo), or
- by dissolution of tungstic acid (when the metal is W) or molybdic acid (when the metal is Mo) ammonia solution, or
- by reaction of W or Mo metal with hydrogen peroxide

According to a further aspect of the invention there is provided a layered inorganic-organic material prepared substantially according to the above method.

20 According to a further aspect of the invention there is provided a method of preparing a layered inorganic-organic material comprising or including

- One or more layers of metal oxide consisting of one or more atomic planes of metal oxide having substantially the ZMO_3 perovskite structure, or derivatives/analogues thereof (M=metal) with the Z sites vacant, and
- one or more layers of organic molecules,

wherein one or more metal-oxide layers alternate with one or more organic layers alternate to form a periodic planar structure,
 and wherein the spacing and electronic coupling between adjacent but separation inorganic layers can be preselected by choice of appropriate organic molecule.

30 According to a further aspect of the invention there is provided a layered inorganic-organic material prepared substantially according to the above methods.

DETAILED DESCRIPTION OF THE INVENTION

The present invention describes novel self-assembling organic/inorganic-oxide multilayer materials based on single or multiple atomic layers of tungsten oxide, molybdenum oxide or other metal oxides interspersed between organic spacer layers.

5 These provide "plastic-like" materials with highly tuneable electronic properties suitable for electronic applications and especially flexible displays, electrochromic devices, sensors and logic and storage devices. In addition the two-dimensional character of these hybrid multilayers offers the possibility of a broad range of other applications including thermoelectric coolers and possibly novel superconductors. The invention also
 10 encompasses other similar organic/ inorganic multilayer materials based on copper, nickel, manganese or ruthenium oxides which display exotic quasi-two-dimensional magnetic properties and superconductivity. The simplicity of their self-assembly, using solution-based fabrication, provides for a cost-advantage technology with all the benefits of plastics, especially flexibility and ease of manufacture.

15 The present invention, by way of example, is directed to the fabrication of materials comprising two-dimensional single atomic layers of corner-shared WO_6 octahedra interspersed between organic spacer layers. Such a material offers the high electron mobility associated with the extended WO_6 sheets together with the flexibility associated with the organic spacer layers which facilitate the self-assembly of the hybrid materials. Such atomically layered materials further provide enhanced and exotic
 20 electronic properties associated with the low dimensionality of the WO_6 sheets.

When a single atomic layer of WO_6 octahedra is formed then its chemical formula is WO_4^{2-} (the four in-plane oxygens all being shared by adjacent octahedra while the two apical oxygens are confined wholly to a single octahedron). Thus two additional cation
 25 charges need to be supplied per formula unit to stabilise the structure. In the present invention these additional charges (providing the charge balance) are introduced using ammonium ions positioned in the vacant A-sites above and below the WO_4^{2-} layer. These ammonium ions may terminate each end of an organic molecule as illustrated schematically in Figure 2 with the structure repeated indefinitely in the c-direction
 30 providing an organic/inorganic multilayered hybrid material where the organic spacer layer is a diammonium organic cation. Such a structure possesses highly anisotropic properties and strong two-dimensional in-plane electronic properties. The organic spacer molecules

may be, for instance, a diammoniumalkane cation generally represented as $\text{NH}_3(\text{CH}_2)_n\text{NH}_3^{2+}$ and the hybrid compound may be synthesized for example by reacting a diaminoalkane with tungstic acid in solution, viz:



We may read Mo (molybdenum) or Vanadium for W (tungsten) in all of the above considerations or these may be used in combination. Furthermore other high valency metals such as Ti, Nb, Ta, Ru and Re may be used in partial combination with Mo, W or V. Other known diammonium organic ions will also provide possible stabilising organic spacer molecules. For synthesis in solution tungstic (or molybdic) acid may be dissolved in ammonia solution or, alternatively, the metals W or Mo may be reacted with hydrogen peroxide to form peroxy polytungstate or peroxy polymolybdate acids as precursors. Other standard solution precursors for the metal oxide species will be known in the art.

15 Further novel compounds of the invention include higher-order organic/inorganic-oxide materials. These comprise multilayers of two or more adjacent sheets of WO_6 or MoO_6 octahedra separated by the organic spacer layer. Such materials are of general formula $\text{NH}_3\text{ANH}_3\cdot\text{W}_m\text{O}_{3m+1}$ where $m=1, 2, 3$, or, more specifically in the case of a diammoniumalkane spacer molecule, $\text{NH}_3(\text{CH}_2)_n\text{NH}_3\cdot\text{W}_m\text{O}_{3m+1}$ where $m=1, 2, 3$. A two-layer compound of this sort is shown schematically in Figure 3.

The novel material structures shown in Figures 2 and 3 may be described as *eclipsed* structures where the octahedra are aligned along the c-axis each in one layer positioned directly above those in adjacent layers across the organic layer. The organic molecules in such materials as described above have two active radical sites (two ammonium groups) and are referred to as bidentate. The invention also comprises a class of *staggered* organic/inorganic-oxide hybrid materials where the octahedra in adjacent layers are staggered in alignment. Here the organic spacer molecules may be monoammonium organic molecules, $\text{A}.\text{NH}_3^+$, which contain a single ammonium group and may be referred to as monodentate. These may stack in interdigitated configuration, with weak bonding between the organic molecules, to provide staggered oxide layers as shown in Figure 4.

The monodentate organic ammonium ions in these structures may comprise an alkyl ammonium ion, $\text{CH}_3(\text{CH}_2)_m\text{NH}_3^+$ with $m=1, 2, 3, \dots 12$, or an aromatic ammonium ion such as benzyl ammonium ion $\text{C}_6\text{H}_5\text{CH}_2\text{NH}_3^+$ or phenethyl ammonium ion $\text{C}_6\text{H}_5\text{CH}_2\text{CH}_2\text{NH}_3^+$. Again higher-order structures, of general formula $(\text{A}.\text{NH}_3)_2\text{W}_n\text{O}_{3n+1}$,

5 are included in the invention in which two or more oxide layers may be interspersed between the organic spacer layers. A schematic example of the $n=2$, double layer compound is shown in Figure 5. The organic/inorganic-oxide hybrid compounds may be synthesized from solution using tungstic acid (or molydbic acid) and benzyl amine or phenethylamine.

10 It can be seen that these novel organic/inorganic-oxide compounds provide *great structural diversity* which allows considerable control of the coupling between the oxide layers and hence the degree of anisotropy and the effective dimensionality of the electronic states. Moreover, as noted they may be constructed using W, Mo, V or any combination thereof and in addition other metallic cations such as Ti, Nb, Ta, Re and Ru, or other 15 suitable metal, may be incorporated to stabilise the structures and enhance their chemical stability, as well as dope the oxide layers to provide semiconducting or metallic properties. Electronic doping may also be achieved through incorporation of organic dopants such as methyl ammonium ions. They may be doped by intercalating alkali ions, by electrochemical redox of the organic spacer molecules or by field-effect insertion of 20 carriers in a three-terminal device such as a field-effect transistor, as is known in the art.

Further novel cuprate, nickelate, ruthenate and manganese-based organic/inorganic hybrids are also included in the invention. The basic structure of the materials described above is contingent upon the octahedral oxygen coordination of a hexavalent cation such as W^{6+} or Mo^{6+} . Organic/inorganic-oxide structures may be 25 fabricated with divalent cations such as Cu^{2+} and Ni^{2+} where these are coordinated into a corner-shared square-planar structure (i.e. with no apical oxygen) or with tetravalent cations such as Mn^{4+} or Ru^{4+} where these are coordinated into a corner-shared square-pyramid structure. In both cases higher order structures with two or more oxide layers each separated by an alkali earth ion, preferably calcium, situated in the perovskite A-site, may 30 be fabricated. The general formulae for these additional novel materials are $\text{NH}_3.\text{A}.\text{NH}_3\text{CuO}_2$, $(\text{A}.\text{NH}_3)_2\text{CuO}_2$, $\text{NH}_3.\text{A}.\text{NH}_3\text{Ca}_{m-1}\text{Cu}_m\text{O}_{2m}$, $m=1, 2, 3, \dots$, $(\text{A}.\text{NH}_3)_2\text{Ca}_{m-1}\text{Cu}_m\text{O}_{2m}$, $m=1, 2, 3, \dots$, $\text{NH}_3.\text{A}.\text{NH}_3\text{NiO}_2$, $(\text{A}.\text{NH}_3)_2\text{NiO}_2$, $\text{NH}_3.\text{A}.\text{NH}_3\text{Ca}_{m-1}\text{Ni}_m\text{O}_{2m}$, $m=1,$

2, 3,..., $(A.NH_3)_2Ca_{m-1}Ni_mO_{2m}$, m=1, 2, 3,..., and $NH_3.A.NH_3MnO_3$, $(A.NH_3)_2MnO_3$, $NH_3.A.NH_3Ca_{m-1}Mn_mO_{2m+2}$, m=1, 2, 3,..., $(A.NH_3)_2Ca_{m-1}Mn_mO_{2m+2}$, m=1, 2, 3,..., $NH_3.A.NH_3RuO_3$, $(A.NH_3)_2RuO_3$, $NH_3.A.NH_3Ca_{m-1}Ru_mO_{2m+2}$, m=1, 2, 3,..., $(A.NH_3)_2Ca_{m-1}Ru_mO_{2m+2}$, m=1, 2, 3.... These may be doped with carriers in the same ways as described above. The wide variety of these materials provide for different types of magnetic ordering (e.g. just by altering the structure from staggered to eclipsed) and other correlated states such as charge density waves and low-dimensional superconductivity. The discovery of surface superconductivity at 91K on the surface of lightly doped Na_xWO_3 single crystals [see Y. Levi et al., " *Europhys. Lett.* 51, 564 (2000)] raises the prospect of high temperature superconductivity in the present novel two-dimensional compounds.

Further monodentate molecules which may be used in the organic spacer layer include biphenylamine, $C_6H_5.C_6H_4.NH_2$, thiophenamine, $C_4SH_3.NH_2$, bithiophenamine, $C_4SH_3.C_4SH_2.NH_2$, and trithiophenamine, $C_4SH_3.(C_4SH_2)_2NH_2$. Further bidentate molecules which may be used in the organic spacer layer include biphenyldiamine, $NH_2.(C_6H_4)_2.NH_2$, diaminothiophene, $NH_2.C_4SH_2.NH_2$, diaminobithiophene, $NH_2.(C_4SH_2)_2.NH_2$, and diaminotri thiophene, $NH_2.(C_4SH_2)_3.NH_2$. These and other conjugated molecules offer the possibility of conducting organic spacer layers and also the possibility of polymerisation of the organic spacer layer. Control of the interlayer coupling between inorganic oxide layers using conducting organic molecules provides a key means to control the phase behaviour and electronic properties of the novel oxides.

The compounds of the invention may be incorporated into devices by the production of films using methods known in the art such as spin coating, dip coating, spraying, painting, printing and such like.

Their great utility is that they self assemble by precipitation out of solution. They may then be dissolved in organic solvents and spin-coated onto suitable substrates and for example onto single-crystal perovskite substrates, such as $SrTiO_3$, with matching lattice parameters. The substrate lattice parameter may be selected to stabilise the films. Through careful heat treatment c-axis oriented films may be fabricated. Devices such as sensors, electrochromic films and thin-film field-effect transistors may be fabricated using these methods. Field-effect transistors may be made using the organic/inorganic-oxide as the semiconducting channel and device characteristics may be optimised by control of doping

and of the spacing of the oxide layers. Sensor applications may be implemented based on the change in doping state induced by adsorbed gases.

A further means of assembling the present organic/inorganic oxide layered materials is to grow them onto self-assembled monolayers which are terminated by amine groups. It is well known that thiol-, isocyanide- or thioacetyl-terminated organic molecules self assemble onto gold substrates. If these molecules are terminated at the opposing end by an amine group then these molecules may be assembled as a monolayer onto gold then the organic/inorganic oxide material may be assembled onto the monolayer. Such amino-thiolated organic molecules for fabricating self-assembled monolayers include the 1-n 5 aminoalkanethiols, HS.(CH₂)_n.NH₂ n=1, 2, 3,..., aminophenylthiol HS.C₆H₄.NH₂, aminobiphenylthiol HS.(C₆H₄)₂.NH₂, amino-bithiophene-thiol HS.(C₄SH₂)₂.NH₂, and other 10 related compounds. Variations of these, including the separation of amine or thiol groups from the phenyl groups using alkane chains, and the replacement of thiol groups by cynaide groups or thioacetyl groups in all of the above will be obvious extensions within 15 the scope of the present invention.

BRIEF DESCRIPTION OF THE FIGURES

The invention is described with reference to the Figures in which:

- Figure 1:** illustrates a schematic plan view of the structure of WO₃ or MoO₃;
- Figure 2:** illustrates a schematic side cross section view of one material in accordance with the invention;
- Figure 3:** illustrates a schematic side cross section view of a second material in accordance with the invention;
- Figure 4:** illustrates a schematic side cross section view of a further material in accordance with the invention;
- Figure 5:** illustrates a schematic side cross section view of a further material in accordance with the invention;
- Figure 6:** illustrates an XRD pattern of one material of the invention;
- Figure 7:** is a plot of c-axis spacing vs alkane number;
- Figure 8:** is a plot of c-axis spacing vs alkane number;
- Figure 9:** is a scanning electron micrograph of a material of the invention;
- Figure 10:** is a scanning electron micrograph of a material of the invention.

With specific reference to the figures:

Figure 1 shows a schematic diagram of a plan view of the structure of WO_3 (or MoO_3) in which the corner-shared WO_6 octahedra are represented by the shaded corner-shared diamonds, 1. Doping is achieved by substituting cations into the vacant 12-coordinated perovskite A-sites.

Figure 2 shows a schematic diagram of a cross sectional side view of the structure of the hybrid organic/ inorganic multilayered compound $\text{NH}_3\text{ANH}_3\cdot\text{WO}_4$ where $\text{NH}_3\text{ANH}_3^{2+}$ may be a diammoniumalkane chain. The single atomic layer of WO_6 octahedra, 1 are again represented by the shaded diamonds and the NH_3^+ groups, 2 are represented by the circles at either end of the alkane chains, with the organic group A is represented by the ellipses.

Figure 3 shows a schematic diagram of a cross sectional side view of the structure of the two-layer hybrid organic/ inorganic multilayered compound $\text{NH}_3\text{ANH}_3\cdot\text{W}_2\text{O}_7$ where $\text{NH}_3\text{ANH}_3^{2+}$ may be a diammoniumalkane chain. The single atomic layer of WO_6 octahedra, 1 are again represented by the shaded diamonds and the NH_3^+ groups, 2 are represented by the circles at either end of the alkane chains, with the organic group A is represented by the ellipses. In general, any number of oxide layers may be interspersed between the organic spacer layers providing the general formula $\text{NH}_3\text{ANH}_3\cdot\text{W}_m\text{O}_{3m+1}$.

Figure 4 shows a schematic diagram of a cross sectional side view of the structure of the staggered hybrid organic/ inorganic-oxide multilayered compound $(\text{ANH}_3)_2\text{WO}_4$ where A is an organic molecule represented by the ellipse. The WO_6 octahedra, 1 are again represented by the shaded corner-shared diamonds and the NH_3^+ groups, 2 by the circles.

Figure 5 shows a schematic diagram of the structure of the 2nd member of the general staggered structural series $(\text{A.NH}_3)_2\text{W}_n\text{O}_{3n+1}$. The double layer of W_2O_7 is shown by the double layer of shaded corner-shared diamonds and the interdigitated ellipses A represent alkyl or aromatic organic groups terminated by ammonium ions, 2 (circles).

Figure 6 shows the x-ray diffraction pattern for diammonium-hexane-tungstate as described in example 2.

Figure 7 shows a plot of c-axis spacing as a function of alkane number, n, for diammonium-alkane-tungstate prepared at pH 10 for n=2, 6 and 12 as described in examples 2, 3 and 4.

Figure 8 shows a plot of c-axis spacing as a function of alkane number, n, for diammonium-alkane-tungstate prepared at pH \square 10 and pH \square 6 for n=2, 6 and 12 as described in example 5.

Figure 9 shows a scanning electron micrograph of phenylenediammonium-tungstate powder prepared at pH \square 10 as described in example 10.

Figure 10 shows a scanning electron micrograph of phenylenediammonium-tungstate powder prepared at pH \square 1.5 as described in example 10.

EXAMPLES

10 Example 1

A quantity 3.429 g of H₂WO₄ was weighed out and dissolved in 60 ml of ammonia solution with gentle heating to 70°C. After cooling 3 ml of Benzylamine was added and the mixture stirred and heated under flowing nitrogen. The excess ammonia was evaporated off slowly, leaving the wet precipitate, which was then filtered and dried.

15 X-ray Diffraction was performed on the resultant white powder. The powder diffraction pattern revealed a major peaks corresponding to d-spacings of 16.48 Å and 15.59 Å. One of these, possibly the smaller, is consistent with the expected lattice parameter for dibenzylammonium tungstate, (C₆H₅CH₂NH₃)₂WO₄. The other compound is as yet unidentified.

20 A film was made by redissolving some of the powder in anhydrous ethanol, spinning onto a SrTiO₃ single-crystal substrate at 2000 rpm and drying for 15 minutes under flowing nitrogen gas at 80°C.

Example 2

25 Hybrid organic/inorganic oxide materials were synthesized by reacting diaminoalkanes with tungstic acid in stoichiometric proportions. Tungstic acid H₂WO₄ was dissolved in ammonia solution and 1-6 diaminohexane was separately dissolved in ammonia solution. The two solutions were mixed in stoichiometric proportions and the resulting solution, with pH of about 10, was heated at 90°C to evaporate the ammonia. The 30 pH remained about 9.5 once the ammonia had evaporated. The solution was then heated at 80°C to evaporate the remaining water and a powder precipitated. The powder was subjected to x-ray powder diffraction (XRD) and the XRD pattern is shown in Figure 6.

The marked diffraction lines correspond to the 001, 002 and 003 reflections of a layered compound with c-axis lattice parameter $c=12.27\text{ \AA}$. The compound is identified as the novel compound diammoniumhexane-tungstate with formula $(\text{NH}_3)_2(\text{CH}_2)_6\text{WO}_4$.

5 **Example 3.**

The process of example 2 was repeated but using 1-2 diaminoethane dissolved in ammonia solution. The XRD pattern of the resultant powder revealed a c-axis spacing of 7.39 \AA . The compound is identified as the novel compound diammoniummethane-tungstate with formula $(\text{NH}_3)_2(\text{CH}_2)_2\text{WO}_4$.

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Example 4.

The process of example 2 was repeated but using 1-12 diaminododecane dissolved in ethanol. The XRD pattern of the resultant powder revealed a c-axis spacing of 19.59 \AA . The compound is identified as the novel compound diammoniumdodecane-tungstate with formula $(\text{NH}_3)_2(\text{CH}_2)_{12}\text{WO}_4$. The c-axis spacings for the products of examples 2, 3 and 4 are plotted as a function of alkane number in Figure 7. The result is a straight line of equation $c = 1.22n + 4.99\text{ \AA}$. This is consistent with the alkane chains in each compound being aligned perpendicular to the tungstic oxide layers and the nitrogen atoms in the ammonium groups lying in the vacant perovskite Z sites in the plane of the apical oxygen atoms in the ZMO_3 structure. Firstly the additional c-axis spacing (1.22 \AA) for each additional carbon atom in the chain length is consistent with the C-C bond-length of 1.54 \AA and the C-C-C bond angle of 109.28° , giving a C-C spacing projected along the length of the alkane chain of 1.25 \AA . Now consider the offset, d_0 , between the nitrogen atoms in the ammonium groups and the vacant perovskite Z sites in the ZMO_3 structure. The intercept value of 4.99 \AA comprises the W-W bondlength (3.75 \AA) and a projected C-N bondlength (1.22 \AA) plus the offset d_0 , namely $4.99 = 3.75 + 1.22 + d_0$. Clearly d_0 is essentially zero. These compounds therefore form model eclipsed structures comprising a single atomic layer of WO_4 consisting of a corner-shared WO_2 layer capped above and below by apical oxygens completing the octahedral coordination of W. The diammonium alkane chains extend perpendicular to these WO_4 layers with the N atom sited approximately in the plane of the apical oxygen atoms and in the face-centred location as if the NH_3 group were the Z atom in the perovskite ZMO_3 structure. Scanning electron

microscopy shows these materials to exhibit a very flat micaceous crystallites consistent with their layered structure.

Example 5

5 Examples 2, 3 and 4 were repeated with the exception that the pH off he mixed solution was lowered by adding either hydrochloric acid or nitric acid. As described in the examples, at pH about 10 the canonical layered diaminoalkane tungstates were produced. At pH about 8, and for both acids, a purely inorganic oxide precipitated out, that is having no organic component. This was identified using XRD to be ammonium tungstate
 10 $(\text{NH}_4)_{10}\text{W}_{12}\text{O}_{41}$. At pH 6 or lower HCL resulted again in the formation of $(\text{NH}_4)_{10}\text{W}_{12}\text{O}_{41}$ but HNO_3 resulted in powders with a large c-axis spacing and these powders were not single phase. Figure 8 shows the c-spacings obtained in this way at pH=6 for n=2, 6 and 12 plotted as a function of the alkane number n. They yield a straight line fit of $c = 1.22n + 6.96$. These novel materials therefore have a c-spacing which is 2 Å larger than that of the canonical layered organic/inorganic materials. Synchrotron x-ray diffraction reveal these to consist of an ordered array of oval shells of tungsten oxide separated by diammoniumalkane chains. These are not canonical layered organic/inorganic materials which evidently require pH in excess of 8 to form.
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Example 6

20 Examples 2, 3 and 4 were repeated using molybdic acid, in the form of ammonium dimolybdate, instead of tungstic acid. This resulted at high pH in the same canonical layered organic inorganic materials of chemical formula $(\text{NH}_3)_2(\text{CH}_2)_n\text{WO}_4$ for n=2, 6 and 12. Scanning electron microscopy showed these also to have flat crystallites.
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Example 7

Benzylamine and tungstic acid were reacted in stoichiometric proportions as described in example 2. After evaporation of ammonia and water a white powder was obtained. This showed a single-phase layered structure with lattice c-parameter of 16.5 Å.
 30 This is a canonical layered organic/inorganic oxide of formula $(\text{C}_6\text{H}_5\text{CH}_2\text{NH}_3)_2\text{WO}_4$.

Example 8

Analine and tungstic acid were reacted in stoichiometric proportions as described in example 2 at pH≈10. After evaporation of ammonia and water a white powder was obtained. This revealed ammonium tungstate, $(\text{NH}_4)_{10}\text{W}_{12}\text{O}_{41}$, as the dominant phase and the absence of any canonical layered organic/inorganic oxide. The example was repeated at pH=6 with the same result. The example was repeated using molybdic acid with the same result.

Example 9

Phenethylamine, $\text{C}_6\text{H}_5(\text{CH}_2)_2\text{NH}_2$ and tungstic acid were reacted in stoichiometric proportions as described in example 2 at pH≈10. After evaporation of ammonia and water a white powder was obtained. This showed a single-phase layered structure. This is a canonical layered organic/inorganic oxide phenethylammonium tungstate of formula $(\text{C}_6\text{H}_5(\text{CH}_2)_2\text{NH}_3)_2\text{WO}_4$.

Example 10

Phenylenediamine, $\text{NH}_2\text{C}_6\text{H}_4\text{NH}_2$ and tungstic acid were reacted in stoichiometric proportions as described in example 2 at pH≈10. After evaporation of ammonia and water a black powder was obtained. XRD of this powder revealed a layered structure with lattice d-spacings of 14.2Å and 12.8Å. The scanning electron micrograph shown in Figure 9 indicates crystallites in the form of flat sheets consistent with their layered structure. The example was repeated at pH≈1.5. The resultant crystallites were very flat and platey as shown in Figure 10 with d-spacing 12.6Å.

Example 11

Example 10 was repeated using molybdic acid in the form of ammonium dimolybdate. The powder precipitated at pH≈10 was examined by visible and UV spectroscopy. This revealed an apparent plasma edge at 1.5eV consistent with the formation of conductive phenylenediammonium links between the oxide layers. The example was repeated at pH≈1.5. Flat crystallites were obtained with d-spacing ≈12.8Å.

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